

## PERIODIC CLASSIFICATION - 1

STD XII  
CODE: AX

MARKS: 50  
TIME: 1Hr

### UNIT TEST PART - I

CHOOSE THE CORRECT ANSWER:

(10x1=10)

- Modern periodic law is based on  
(a) Atomic mass (b) atomic number (c) periodicity (d) rows and columns
- The eka-aluminium as per Mendeleev's classification is  
(a) Silicon (b) germanium (c) boron (d) gallium
- The IUPAC name of element with atomic number 118 is  
(a) Ununoctium (b) Ununennium (c) Unbinilium (d) Ununnilium
- Elements of group 17 are called  
(a) Noble gases (b) chalcogens (c) halogens (d) alkalimetals
- The no of elements in the fourth period is  
(a) 2 (b) 8 (c) 18 (d) 32
- The general outermost electronic configuration of p - block elements is  
(a)  $ns^2 np^{1-6}$  (b)  $ns^{1-2}$  (c)  $(n-1)d^{1-10} ns^{1-2}$  (d)  $ns^2 np^3$ .
- Main group elements consists of  
(a) s-block and p-block (b) d-block (c) s-block and d-block (d) s-block and f-block
- Which is a metalloid?  
(a) Li (b) Ca (c) Ge (d) Cl
- Which one among the following species has the largest atomic radius?  
(a) Na (b) Mg (c) Al (d) Si
- Which of the following has highest ionisation potential?  
(a) Na (b) Mg (c) C (d) F

### PART - II

ANSWER ALL THE QUESTIONS:

(10X3=30)

- What is Dobereiner's triad? Give an example.
- State Mendeleev's periodic law.
- Write note on gaps in the Mendeleev's periodic table
- How does atomic radius varies in a period and group.
- Define ionization potential.
- Boron has lower ionization energy than Beryllium. Why?
- Electron affinity of fluorine is less than that of chlorine. Why?
- Define second electron gain enthalpy.
- Define electro negativity.
- Why are d-block elements called transition elements?

### PART III

- (a) Write characteristics of Mendeleev's periodic table
- (b) Describe the factors influencing ionization energy.